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3. Convert the temperature to K 4. Now plug this in the formula 5. This would be a negative value because heat flows from the system.  $\Delta S = q_{rev} / T$ . Chemical Thermodynamics Example 9.2 The element mercury, Hg, is a silvery liquid at room temperature. The normal freezing point of ... (1 / 6.01 10 / ) 3 310 19.4 / rev

### **Chapter 19 Chemical Thermodynamics**

5. Calculate the mole fraction, molarity and molality of NH<sub>3</sub> if it is in a solution composed of 30.6 g NH<sub>3</sub> in 81.3 g of H<sub>2</sub>O. The density of the solution is 0.982 g/mL and the density of water is 1.00 g/mL. Molarity: 15.8 M NH<sub>3</sub>, molality: 22.1 molal NH<sub>3</sub>, mole fraction(NH<sub>3</sub>): 0.285

### **Practice Problems: Solutions (Answer Key)**

Source: 3-4 Practice Problems, Chemistry Connections supplement (Prentice-Hall) Alpha & Beta Decay 1. Write a nuclear equation for the alpha decay of <sup>231</sup>Pa.  ${}_{91}^{231}\text{Pa} \rightarrow {}_{89}^{227}\text{Ac} + {}_2^4\text{He}$  2. Write a nuclear equation for the beta decay of <sup>223</sup>Fr.  ${}_{87}^{223}\text{Fr} \rightarrow {}_{88}^{223}\text{Ra} + {}_{-1}^0\text{e}$  3. Write a

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nuclear equation for the alpha decay of  $^{149}\text{Sm}$ .  $^{62}\text{Sm}$  ...

### Chemistry 11 3 Practice Problems Answers

3—3 Practice Problems (continued) 15. Write the chemical symbol for the ion with 21. How many protons, neutrons, and 29 protons and 27 electrons, electrons are present in the  $\text{n}4$  ion? 16. How many protons, neutrons, and 22. Write the complete chemical symbol for electrons are present in the  $\text{Ni}2$  ion? the ion with 84 protons, 125 neutrons, and ...

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0.710 kg of sodium carbonate (washing soda),  $\text{Na}_2\text{CO}_3$ , in 10.0 kg of water—a saturated solution at  $0^\circ\text{C}$ ; 125 g of  $\text{NH}_4\text{NO}_3$  in 275 g of water—a mixture used to make an instant ice pack; 25 g of  $\text{Cl}_2$  in 125 g of dichloromethane,  $\text{CH}_2\text{Cl}_2$ ; 0.372 g of histamine,  $\text{C}_5\text{H}_9\text{N}$ , in 125 g of chloroform,  $\text{CHCl}_3$

### 8.3: Concentrations of Solutions (Problems) - Chemistry ...

Example #13: A golden-colored cube is handed to you. The person wants you to buy it for \$100,

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saying that is a gold nugget. You pull out your chemistry text and look up gold's density and read that its density is  $19.32 \text{ g/cm}^3$ . You measure the cube and find that it is 2.00 cm on each side, and weighs 40.0 g.

### **ChemTeam: Density**

Name \_\_\_\_\_ Date \_\_\_\_\_ 19—1 Practice Problems 1. 2. 3. 4. 5. 8 What is the concentration of  $\text{OH}^-$  ions in saturated limewater if  $[\text{H}^+] = 3.98 \times 10^{-11} \text{ M}$ ? Is acidic, basic, or neutral?

### **Kennedy HS**

Chemistry 10th Edition Whitten, Kenneth W.; Davis, Raymond E.; Peck, Larry; Stanley, George G. Publisher Brooks/Cole Publishing Co. ISBN 978-1-13361-066-3

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### **Chem210 General Chemistry II Practice Tests**

Practice Problem 5. What is the value of a gold ingot 20.0 cm long by 8.5 cm wide by 6.0 cm tall, if the mass of a cubic centimeter of gold is 19.3 grams and the price of gold is \$356 per ounce? Click here to check your answer to Practice Problem 5

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