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Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

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**Chapter 10 Chemical
Quantities Chapter
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CHAPTER 10: Chemical

Quantities BASICS: •

The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he

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decided to use the

**CHAPTER 10:
Chemical Quantities**

Chapter 10 Chemical
Quantities 241 Section
Review Objectives •

Relate Avogadro's
number to a mole of a
substance • Calculate
the mass of a mole of
any substance •

Describe methods of
measuring the amount
of something •

Compare and contrast
the atomic mass of an

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element and its molar
mass Vocabulary Key
Equations • moles
representative
particles •
representative
particles moles.

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Chapter 10 Chemical
Quantities91 SECTION
10.1 THE MOLE: A
MEASUREMENT OF
MATTER (pages
287-296) This section
defines the mole and
explains how the mole
is used to measure
matter. It also teaches

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Answer Key

you how to calculate
the mass of a mole of
any substance.
Measuring Matter
(pages 287-289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

Chapter 10 "Chemical
Quantities" Vocab. the
SI unit representing
 6.02×10^{23}
representative
particles of a

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substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

Quia - Chapter 10 **"Chemical** **Quantities" Vocab**

Chapter 10 (Chemical Quantities) Test Study Guide The mole is the

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SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

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CHAPTER 10: Chemical
Quantities - Scarsdale
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Schools ... The basic unit that is used to determine the amount of a chemical

substance is called ...

Chapter Review ... This calculation will get you the answer of 2.08×10^{-3} ...

Chapter 10 Chemical Quantities Review Answers

CHAPTER 10: Chemical Quantities. CHAPTER 10: Chemical Quantities BASICS: •

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The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

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Answers Chapter 10

... Quantities Test A

Chapter 10 Chemical
Quantities 243 Section

Review Objectives •

Convert the mass of a
substance to the

number of moles of a
substance, and the

number of moles of a
substance to mass •

Calculate the volume
of a quantity of gas at

STP Vocabulary •

Avogadro's hypothesis

• standard

temperature and

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Quantities
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pressure (STP) • molar
volume Key Equations
• mass (grams)
number of moles

Chapter 10 Chemical Quantities

Interpreting

Graphics Answers

the SI unit representing

6.02×10^{23}

representative

particles of a

substance: Avogadro's

number: 6.02×10^{23}

particles: standard

temperature and

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pressure (00 C, 1 atm)
the temperature and
pressure at which one
mole of gas occupies a
volume of 22.4 L:
molar volume: volume
of a gas that contains
one mole of the gas, is
22.4 L at STP:
Avogadro ...

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"Chemical Quantities"

chapter 2/chapter
10/chapter 11 chemical
quantities. day: class

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activity: assignment: w
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s: wed. 11/20. go over
test

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world with arduino
make projects

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