

## Chapter 7 Chemical Formulas Compounds Answer Key

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### Chapter 7 Chemical Formulas Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided.

1. Assign the oxidation number to the specified element in each of the following examples: 4 a. S in  $\text{H}_2\text{SO}_3$  6 b. S in  $\text{MgSO}_4$  2 c. S in  $\text{K}_2\text{S}$  1 d. Cu in  $\text{Cu}_2\text{S}$  6 e. Cr in  $\text{Na}_2\text{CrO}_4$  5 f. N in  $\text{HNO}_3$  4 g. C in  $(\text{HCO}_3)_3$  3 h. N in  $(\text{NH}_4)_2\text{SO}_4$  2.  $\text{SO}_2$  2 a.

### 7 Chemical Formulas and Chemical Compounds

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1) determine the mass of the water driven off. 2) determine the mass of the anhydrous salt. 3) determine the percent of water. 4) determine number of mol of water driven off. 5) determine number of mol of anhydrous salt. 6) determine the ratio (each number divided by the smallest. 1) 15.00g - 12.79g = 2.21g. 2) 12.79g.

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The compound's formula must therefore indicate one  $\text{Mg}^{2+}$  cation and two  $\text{Br}^-$  anions. The symbol for the cation is written first. Ions combined:  $\text{Mg}^{2+}$ ,  $\text{Br}^-$ ,  $\text{Br}$  Chemical formula:  $\text{MgBr}_2$  Note that the 2 in  $\text{Br}_2$  is written as a subscript. The charges of the ions are not included in the formula. This is usually the case when writing

### CHAPTER 7 Chemical Formulas and Chemical Compounds

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### **Chemistry Chapter 7 Chemical Formulas and Compounds ...**

Modern Chemistry Chapter 7 Chemical Formulas and Chemical Compounds. monatomic ion. binary compounds. nomenclature. oxyanions. an ion formed from a single atom. (remember to change nonmetal.... compounds composed of two different elements; (compounds of 2.... naming system. polyatomic ions that contain oxygen.

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WRITING BMCS FORMULAS PRACTICE • Write the formulas for the following compounds: • 1. oxygen difluoride • 2. carbon tetraiodide • 3. phosphorus trichloride • 4. dinitrogen monoxide • 5. disilicon hexabromide • 6. diphosphorus trioxide • 7. disulfur dichloride

### **CHAPTER 7 CHEMICAL COMPOUNDS & CHEMICAL FORMULAS**

CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS 203 SECTION 7-1 OBJECTIVES Explain the significance of a chemical formula. Determine the formula of an ionic compound formed between two given ions. Name an ionic compound given its formula. Using prefixes, name a binary molecular compound from its formula. Write the formula of a binary molecular compound given its name. C

### **CHAPTER 7 Chemical Formulas and Chemical Compounds**

Chapter 7: Chemical Formulas and Chemical Compounds Section 1: Chemical Names and Formulas Overview We will explain the significance of a chemical formula. We will again look at how to determine the formula unit of an ionic compound. We will name an ionic compound based on its formula unit. We will use prefixes to name binary molecular compounds based on its formula.

### **Chapter 7: Chemical Formulas and Chemical compounds**

Chapter 7: Chemical Formulas and Chemical Compounds Section 7-1: Chemical Names and Formulas 7-1-1 Explain the significance of a chemical formula. The subscript indicates that there are 8 carbon atoms in the molecule of octane. The subscript indicates that there are 18 hydrogen atoms in the molecule of octane. Subscript 2 refers to 2 aluminum atoms.

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Chapter 7 Notes – Chemistry; Chemical Formulas and Chemical Compounds Molecules and Molecular Compounds • A molecule consists of two or more atoms bounded together. Molecules and Chemical Formulas • Each molecule has a chemical formula. • The chemical formula indicates: 1. Which atoms are found in the molecule, and 2.

### **Chapter 7 Notes - Chemistry; Chemical Formulas and ...**

In Chapter 7, students will learn how to write the chemical formula of any given compound, either ionic or molecular, given that compound's scientific name (or vice versa). Students will...

### **Chapter 7 - Chemical Formulas & Chemical Compounds - yazvac**

Determining the formula unit starting from the name Example: Copper (II) chloride 1. Write the ions, cation first 2. Criss-cross 3. Check charges to

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make sure they balance 4. Write formula unit  $\text{Cu}_2\text{Cl}_2$

### Chapter 7, Sections 1-2 - [doralacademyprep.enschool.org](http://doralacademyprep.enschool.org)

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### Chapter 7 Chemical Formulas | [wrobelchemistry](http://wrobelchemistry.com)

CHAPTER 7: CHEMICAL NOMENCLATURE. ~Writing Chemical Formulas. ~Naming Chemical Compounds (Ionic, Covalent, Acids) ~Mathematical Analysis of Chemical Formulas.

### Chapter 7 - Chemical Formulas and Nomenclature

"A chemical formula indicates the \_\_\_\_\_ number of \_\_\_\_\_ of each kind in a chemical compound.", What do subscripts tell you about a chemical compound?, The polyatomic ion chlorite ( $\text{ClO}_2^-$ )- bonds with copper to form  $\text{Cu}(\text{ClO}_2)_2$ , a compound called copper (II) chlorite. How many atoms of each element are there in copper chlorite?, Write down the name of the following ion:  $\text{V}^{4+}$ , and determine whether ...

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