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A solution is made so that it is 2.5 M $\text{Ca}(\text{NO}_3)_2$. Therefore the concentration of Ca^{2+} is 2.5 M and the concentration of NO_3^- is 5.0 M. The total concentration of all particles is 7.5 M. Explain. Since each mole of $\text{Ca}(\text{NO}_3)_2$ breaks up into 1 mole of Ca^{2+} , the concentration of Ca^{2+} would be the same as the concentration of $\text{Ca}(\text{NO}_3)_2$.

ChemQuest 39-41 Answers

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Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions

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28. Name: Answer the questions that follow based on this
chemical equation: $\text{Na} + \text{MgCl}_2 \rightarrow \text{NaCl} + \text{Mg}$. a) Why can't NaMg be

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concentration in units of molarity by following these steps: a) Convert the grams of NaCl to moles of NaCl. b) Calculate the litres of solution by dividing the given millilitres by 1000. c) Divide moles by the litres of solution. 4. Calculate the concentration (in units of molarity) of the following solutions: a)

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42.1 g of $\text{Ca}(\text{NO}_3)_2$

Molarity ChemQuest WS

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