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Author: Ted
Gibb, Lucille
Davies, Maurice
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38.00 mL. Solution:

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can be determined: 23
 $2 \text{ HC}_2\text{H}_3\text{O}_2(\text{aq})$ 23
 $2 \text{ HC}_2\text{H}_3\text{O}_2(\text{aq})$ $23 \text{ HC}_2\text{H}_3\text{O}_2$
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Solution: OH OH OH
0.50 mmol mL --
-- = x = n c V x 25.0 mL
OH n - = 12.5 mmol
HCl(aq) HCl(aq)
[HCl(aq)] 12.5 mmol =
= n V 0.10 mmol
HCl(aq) /mL V = 120 mL
2. (a) Given: V HCl(aq)

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= 25.00 mL; [HCl(aq)]

= 0.350 mol/L; V

NaOH(aq) = 5.0 mL;

[NaOH(aq)] = 0.500

mol/L; Required:

[H⁺(aq)], pH Analysis:

Before titration,

amount of H⁺ is: H(aq)

H(aq) H(aq) 0.350

mmol mL ++ += x =

nc V

Section 8.7:
Acid-Base Titration
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